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Intermolecular forces practice worksheet answers

This one page (front/back) worksheet requires students to draw Lewis structures, determine the IMFs present for a particular substance, and compare properties based on IMFs, including boiling point and vapor pressure. Students should be able to rank the strengths of LDF, dipole-dipole, and hydrogen bonding, as well as know that compounds can have stronger LDFs when they have more electrons. Answer key is provided. This is great for extra practice, building confidence with basic skills, or for a sub plan. Which will have the strongest dipole (a) F₂ or HF, (b) CH₃Cl or CH₃Br? Answer (a) HF (b) CH₃Cl Which is more polarizable? (a) Cl₂ or I₂, (b) C₂H₆ or C₁₀H₂₂? Answer (a) I₂ (b) C₁₀H₂₂ Which of the following materials is likely to have (a) no dipole-dipole forces, but the largest London dispersion forces, (b) the largest dipole-dipole intermolecular forces: I₂, He, H₂S, H₂Te. Answer (a) I₂, it is the largest nonpolar molecule. (b) H₂S, S is more electronegative and will make the molecule more polar. What kind of attractive forces must be overcome to (a) sublime Al₂, (b) vaporize C₂H₅F, (c) boil hydrogen fluoride, HF, (d) melt LiBr? Explain. Answer (a) Nonpolar - London Dispersion Forces. (b) Polar - Dipole-dipole attractions (and dispersion forces). (c) Very polar - Hydrogen bonds between the molecules. Hydrogen bonds form when hydrogen is covalently bonded to N, O, or F. (d) No molecules, so there are no intermolecular forces - Ionic bonds. Which of the following materials is likely to have (a) no dipole-dipole forces, but the largest London dispersion forces, (b) the largest dipole-dipole intermolecular forces: I₂, He, H₂S, H₂Te. Answer (a) I₂, it is the largest nonpolar molecule. (b) H₂S, S is more electronegative and will make the molecule more polar. Identify the following crystals (A-E) as ionic, polar-molecular, nonpolar-molecular, covalent (network), or metallic. Explain. There could be more than one or none of the crystal types in the chart. CrystalMelting Point (°C)Boiling Point (°C)Electrical Conductivity Solid Liquid A-83.119.54NoNo B-259.14-252.5NoNo C15353000YesYes D6861330NoYes E-56.6-78.5NoNo F-182.48-164.8NoNo G35504827NoNo Answer A ⇒ good separation between melting and boiling points, relatively high m.p. for molecular, no conductivity - polar-molecular. (HF - polar molecule) B ⇒ close and low melting and boiling points, no conductivity - nonpolar-molecular. (H₂ - nonpolar molecule) C ⇒ high mp and bp, good conductor - metallic. (Fe - metal) D ⇒ high melting and boiling points, conducts as a liquid - ionic. (KI - ionic compound) E ⇒ low and close together mp and bp, nonconductor - nonpolar-molecular. (CO₂ - nonpolar molecule) F ⇒ low and close together mp and bp, nonconductor - nonpolar-molecular. (CH₄ - nonpolar molecule) G ⇒ very high mp and bp, nonconductor - covalent (network). (Diamond - covalent network) Indicate the type of crystal (ionic, polar-molecular, nonpolar-molecular, covalent, or metallic) and predict some of the properties (difference in melting and boiling points, conductivity, etc.) for each of the following substances upon solidification: HF, H₂, Fe, KI, CO₂, CH₄, and diamond. Answer HF - polar molecule, good separation between melting and boiling points, relatively high m.p. for molecular, no conductivity - polar-molecular. H₂ - nonpolar molecule, close and low melting and boiling points, no conductivity - nonpolar-molecular. Fe - metal, high or low mp and bp, good conductor - metallic. KI - ionic compound, high melting and boiling points, conducts as a liquid - ionic. CO₂ - nonpolar molecule, low and close together mp and bp, nonconductor - nonpolar-molecular. CH₄ - nonpolar molecule, low and close together mp and bp, nonconductor - nonpolar-molecular. Diamond ⇒ covalent network, very high mp and bp, nonconductor - covalent (network) CrystalMelting Point (°C)Boiling Point (°C)Electrical Conductivity Solid Liquid HF-83.119.54NoNo H2-259.14-252.5NoNo Fe15353000YesYes KI6861330NoYes CO2-56.6-78.5NoNo CH4-182.48-164.8NoNo Diamond35504827NoNo

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